

CHE 255 Phillips

Due at the beginning of class on Feb 3, 2010

Quiz 2 (16 points)

Name _____

ID# _____

For each of the following molecular formulas, draw valid Lewis structures, formal charges and hybridization of the indicated atoms, and approximate bond angles as indicated. Where resonance structures are possible, draw all resonance structures and the resonance hybrid.

1. C_2H_4 ; C-C-H bond angle; formal charge of C; hybridization of C

2. CH_3CONH_2 ; O-C-N bond angle, formal charge of N; hybridization of N

3. $[CH_3CH_2O]^{-1}$; H-C-O bond angle; formal charge of O; hybridization of O

4. CH_3CN ; C-C-N bond angle; formal charge of Cs; hybridization of Cs